

coupling across, *PCCC*,<sup>8</sup> *POCC*,<sup>9</sup> *POCH*,<sup>10</sup> *CCCH*,<sup>11</sup> and *HCCN*<sup>12</sup> and for four-bond coupling across *HC-COP*.<sup>13-16</sup> Preliminary measurements<sup>17</sup> indicate that such relationships may exist for coupling across *CCCC* and *MCCC* where M is either lead or mercury.

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## Additions and Corrections

**Stable Carbocations. CXXX. Carbon-13 Nuclear Magnetic Resonance Study of Halocarbenium Ions. Degree of Halogen "Back-Donation" and Relative Stability of Halocarbenium Ions** [*J. Amer. Chem. Soc.*, **94**, 3551 (1972)]. By GEORGE A. OLAH,\* Y. K. MO, and Y. HALPERN, Department of Chemistry, Case Western Reserve University, Cleveland, Ohio 44106.

The <sup>13</sup>C nmr shift of the carbocation center in the dimethylfluorocarbenium ion [(CH<sub>3</sub>)<sub>2</sub>C<sup>+</sup>F] was reported, due to a computational error made using the indor method, as -142.7 (from CS<sub>2</sub>). The correct value is -89.1 (as also shown by recent FT measurements).

**Conformational Analysis. XXVI. Conformational Equilibria in 5,5-Disubstituted 1,3-Dioxanes** [*J. Amer. Chem. Soc.*, **94**, 8072 (1972)]. By ERNEST L. ELIEL\* and RUDYARD M. ENANOZA, Department of Chemistry, University of Notre Dame, Notre Dame, Indiana 46556].

Professor Anteunis has informed us that Table I in ref 26, part 2 does not correspond to our Table II inasmuch as it refers to the equilibrium B ⇌ A, not A ⇌ B in Scheme II. Therefore the Δ*G*<sup>o</sup> values for the last five entries in our Table II taken from ref 26 should read -0.19, -0.03, -0.32, -0.38, -0.17 kcal/mol, and the signs of Δ*G*<sup>o</sup> in Table II do agree with those in Table III, entries 2, 5, 6, and 7. Moreover, in the case of 5-methyl-5-phenyl-1,3-dioxane the nmr method of ref 26 does lead to a clear choice as to which isomer predominates, based on the chemical shift of the 5-methyl group below the coalescence temperature.

We have also found that the infrared data reported on page 8076, Scheme III were subject to a calibration error; the correct frequencies are **27**, 3435 cm<sup>-1</sup>; **28**, 3450, 3434 cm<sup>-1</sup>. We are not sure whether the NH absorption in **27** indicates intramolecular H bonding;

rather the double absorption in **28** may be due to cis-trans isomerism of the amide. Configurational assignments are not affected by this uncertainty since they rest on the dipole determination of **19**. We thank Dr. W. F. Bailey for experimental assistance and Dr. Jan Stanek (Prague, Czechoslovakia) for helpful correspondence.

**Excited State Carbonyl Species from the Thermal Decomposition of 3,3-Dibenzyl-1,2-dioxetane** [*J. Amer. Chem. Soc.*, **94**, 9277 (1972)]. By WILLIAM H. RICHARDSON,\* FREDERICK C. MONTGOMERY, and MARY B. YELVINGTON, Department of Chemistry, California State University, San Diego, San Diego, California 92115.

Figure 1, line 4 should read "bibenzyl" rather than "dibenzyl ketone."

**Hydrolysis of Imidate Esters Derived from Weakly Basic Amines. Influences of Structure and pH on the Partitioning of Tetrahedral Intermediates** [*J. Amer. Chem. Soc.*, **95**, 1253 (1973)]. By TADASHI OKUYAMA, TERRY C. PLETCHER, DAVID J. SAHN, and GASTON L. SCHMIR,\* Department of Molecular Biophysics and Biochemistry, Yale University School of Medicine, New Haven, Connecticut 06510.

In Table I, the heading of column 5 should read 10<sup>-7</sup>*k*<sub>7</sub>; the heading of column 7 should read 10<sup>-4</sup>·(*k*<sub>5</sub>/*K*<sub>a</sub>). The uppermost line in Figure 6 should be labeled 10<sup>-5</sup>*k*<sub>7</sub>.

**Stable Carbocations. CXLII. 1-Acenaphthenium Ions** [*J. Amer. Chem. Soc.*, **95**, 3698 (1973)]. By GEORGE A. OLAH,\* GAO LIANG, and PHILIP WESTERMAN, Department of Chemistry, Case Western Reserve University, Cleveland Ohio 44106.

The entries for cation **1a** in Table II should read: C<sub>1</sub>, -6.4; C<sub>2</sub>, 148.7; C<sub>3</sub>, C<sub>4</sub>, C<sub>5</sub>, C<sub>7</sub>, 61.0, 61.0, 58.4, 57.6; C<sub>6</sub>, 30.9; C<sub>8</sub>, 45.6; C<sub>9</sub>, C<sub>10</sub>, C<sub>11</sub>, C<sub>12</sub>, 62.2, 52.1, 49.4, 46.8.

**Equilibrium and Kinetics of Glyconitrile Formation in Aqueous Solution** [*J. Amer. Chem. Soc.*, **95**, 3729 (1973)]. By GORDON SCHLESINGER and STANLEY L. MILLER,\* Department of Chemistry, University of California, San Diego, La Jolla, California 92037.

In Table IV, the entries for benzaldehyde should be  $K_3 = 2.3 \times 10^{-11}$ ,  $K_2 = 0.132$ ,  $k_{-2} = 15.0$ . The preferred name for formaldehyde cyanohydrin is glycolonitrile.

**2,4,6-Trisubstituted Pyridines. Synthesis, Fluorescence, and Scintillator Properties** [*J. Amer. Chem. Soc.*, **95**, 4891 (1973)]. By MARJA DEL CARMEN G. BARRIO, JORGE R. BARRIO, GRAHAM WALKER, ARMANDO NOVELLI, and NELSON J. LEONARD,\* Facultad de Farmacia y Bioquímica, Universidad de Buenos Aires, Buenos Aires, Argentina, and the Roger Adams Laboratory, School of Chemical Sciences, University of Illinois, Urbana, Illinois 61801.

Delete the second sentence of the second paragraph on page 4891 and substitute "The Zecher and Kröhnke procedure using phenacylpyridinium salts and chalcones in the presence of acetic acid-ammonium acetate offers a versatile method of synthesis of substituted pyridines."

To ref 6 should be added: W. Zecher and F. Kröhnke, *ibid.*, **94**, 698 (1961); F. Kröhnke and W. Zecher, *Angew. Chem.*, **74**, 811 (1962).

**Ring Inversion in Dioxene. Comparison of the Barrier Heights by Nuclear Magnetic Resonance and Far-Infrared Measurements** [*J. Amer. Chem. Soc.*, **95**, 5129 (1973)]. By R. H. LARKIN and R. C. LORD,\* Spectroscopy Laboratory and Department of Chemistry, Massachusetts Institute of Technology, Cambridge, Massachusetts 02139.

A calculational error in the equilibrium angle of twist in the dioxene molecule from the nmr data has been pointed out to us by Professor F. B. Malloy, Jr. Correction leads to a value of 26.8° in place of the reported 39.3°. This correction increases the discrepancy with the angle estimated from the infrared data from 9 to 22° and makes clear the need for the warning given that the infrared value is the result of "substantial extrapolation of the potential surface as determined from eigenvalues near the potential minimum."

**Synthesis, Chemistry, and Spectroscopy of Some Tris-(pentahaptocyclopentadienyl)uranium(IV) Alkyl and Aryl Compounds** [*J. Amer. Chem. Soc.*, **95**, 5529 (1973)]. By TOBIN J. MARKS,\* AFIF M. SEYAM, and JOHN R. KOLB, Department of Chemistry, Northwestern University, Evanston, Illinois 60201.

Footnote *b* in Table III should read "Calculated by the method of ref 25 using structural data of ref 26a."

The last sentence in the first paragraph ending in column 2, page 5534 should read, "Thus, for a closely related series of compounds, the above mechanism of spin distribution should give upfield proton shifts for  $a_i < 0$  ( $A_i > 0$ ) and downfield shifts for  $a_i > 0$  ( $A_i < 0$ )."

**Barriers to Internal Rotation in Trineopentylbenzenes. III. Nitro-Substituted Compounds** [*J. Amer. Chem. Soc.*, **95**, 5615 (1973)]. By BERTIL NILSSON,\* PER MARTINSON, KÅRE OLSSON, and ROBERT E. CARTER,\* Organic Chemistry 2, Chemical Center, S-220 07 Lund 7, Sweden, and the Department of Organic Chemistry, University of Goteborg and Chalmers Institute of Technology, S-402 20, Goteborg 5, Sweden.

The positions of the two cyclic diagrams on page 5621 should be interchanged; the first sentence of the second column on page 5621 should read as follows: "If the assignment is reversed, *i.e.*, the larger AB shift assigned to the protons of the 4-methylene group in rotamer A, the sign combinations for the 4- and 6-methylenes on going through an exchange cycle will, of course, be the same, but the line shapes will be different from those given above."

**<sup>31</sup>P Nuclear Magnetic Resonance Spectrum of Pentamethylcyclopentaphosphine. Evidence for a Large Stereochemical Dependence of the <sup>1</sup>J(PP) Spin-Spin Coupling** [*J. Amer. Chem. Soc.*, **95**, 6498 (1973)]. By J. P. ALBRAND, D. GAGNAIRE, and J. B. ROBERT,\* Laboratoire de Chimie Organique Physique, Département de Recherche Fondamentale, Centre d'Etudes Nucléaires de Grenoble, F.38041, Grenoble-Cedex, France.

The phosphorus chemical shift values given in Table I are erroneous. They should read:  $\delta(P_1)$ , -16.9;  $\delta(P_2) \delta(P_3)$ , -16.3;  $\delta(P_4) \delta(P_5)$ , -14.6.

**Molybdenum(IV) in Aqueous Solutions** [*J. Amer. Chem. Soc.*, **95**, 6871 (1973)]. By MICHAEL ARDON\* and ARNOLD PERNICK, Department of Inorganic and Analytical Chemistry, The Hebrew University of Jerusalem, Jerusalem, Israel.

Dr. M. Lamache-Duhameaux has drawn our attention to her work, in collaboration with Professor Souhay and Drs. Cadiot and Viostat in which the existence of stable molybdenum(IV) aquo species has been reported [*C. R. Acad. Sci.*, **262**, 1524 (1966); *Bull. Soc. Chim. Fr.*, 592 (1970)].

**Chemically Produced Excited States. Energy Transfer, Photochemical Reactions, and Light Emission** [*J. Amer. Chem. Soc.*, **95**, 7050 (1973)]. By EMIL H. WHITE,\* PETER D. WILDES, JACEK WIECKO, HAROLD DOSHAN, and C. C. WEI, Department of Chemistry, The Johns Hopkins University, Baltimore, Maryland 21218.

In ref 10 the name W. Adam should be added as the terminal author name.

On page 7056, second column, under the heading 3-Methyl-3-phenyl-1,2-dioxetane, the nmr values that are listed are wrong. They should read  $\tau$  5.0 (d) and a singlet at 8.1.